

FARKAS, L. (Budapest); PALLOS, L. (Budapest)

Natural polyhydroxy-benzalcoumaranones (aurones) and their glucosides.
Periodica polytechn chem 4 no.1:73-76 '60. (EEAI 9:12)

1. Organisch-Chemisches Institut der Technischen Universitat,
Budapest.

(Glycosides) (Aurones)
(Benzylidenehydroxybenzofuranone)

HUNGARY/Organic Chemistry. Synthetic Organic Chemistry.

G-2

Abs Jour: Ref Zhur-Khim., No 13, 1958, 43315.

Author : Zemplen Geza, Mester Laszlo, Pallos Laszlo.

Inst :

Title : Partial Acylation of Physetol and Omega-Hydroxy-phloracetophenone. II. Mesylation.

Orig Pub: Magyar tud. akad. Kem. tud. oszt. kozl., 1957, 8.
No 2-3, 377-379

Abstract: 2,4-(OSO₂CH₃)₂C₆H₃COCH₂OH (I) and 2,4,6-(OSC₆H₄CH₃)₂C₆H₃COCH₂OH (II) were obtained. To 2 g omega-acetylphysetol in 20 ml pyridine were added at -5° 2.29 g mesyl chloride. After standing for 50 hours in the cold and out of contact with air, poured on ice; after about 24 hours omega-acetyl-I (III) was obtained, yield 88 %, MP 111° (from alcohol)

Card : 1/2

FARKAS, Lorand; PALLOS, Laszlo

Final structure and synthesis of coreopsin. Magy kem
folyoir 65 no. 7:278-280 Jl '59.

1. Budapesti Muszaki Egyetem Szerves-Kemiai Tanszeke.

PARIS, France - April 19, 1986 - Interview with Valéry J. Ziegler

French Foreign Minister, Paris, France - Lt. Col. Major General Gén. de l'Air
Valéry J. Ziegler

French Foreign Minister, Paris, France - Lt. Col. Major General Gén. de l'Air
Valéry J. Ziegler

FARKAS, Lorand; PALLÓS, László; FÁAL, Zoltán

Final synthesis and composition of sulfurein. Magy kem folyoirat
no. 4:121-123 Ap'60.

1. Budapesti Műszaki Egyetem Szerves-Kémiai Tanszéke.

FARKAS, Lorand; PALLOS, Laszlo

Synthesis of palasitrin, a glucoside of Butea frondosa.
Magy kem folyoir 66 no. 10:391-392 G '60.

1. Budapesti Műszaki Egyetem Szerves Kemiai Tanszeke.

PARKAS, Lorand; PALLOS, Laszlo; HIDASI, Gyorgy

Synthesis of cernuoid and aureusidin. Magy kem folyoir 67 no.9:
388-390 S '61.

1.Budapesti Muszaki Egyetem Szerves Kemial Tanszeke.

PALLOS, Laszlo (Budapest, XI., Gellert ter 4); FARKAS, Lajos, dr. (Budapest, XI., Gellert ter 4).

Aldoses and their glycosides. Pt. 5. periodica polytechn chem
8 no. 3;133-139 '54.

Chair of organic chemistry of Budapest Technical University,
and United Food and Nutriment factory, Budapest. Submitted
February 6, 1954.

L 01186-66

ACCESSION NR: AP5025819

2005/05/07 10:02:02 2012/02/20/02/22

18

B

AUTHOR: Parkas, Lorand; Pallos, Laszlo; Nagyosi, Mihaly

TITLE: Aurones and aurone glucosides. Part 8: Synthesis of maritimine, one of the glucosides in Coreopsis maritima

SOURCE: Magyar kozlai folyoirat. v. 71, no. 6, 1965, 270-272

TOPIC TAGS: carbohydrate, plant chemistry, organic synthetic process

ABSTRACT: [Part 7 of this series was published in *Planta Medica*, Vol 12, 1964, p 296] 6,7-dihydroxy-2-(3,4-dihydroxybenzyl)-chromene-(3)- β -D-glucoside-(6), proven to be identical with one of the glucosides in *Coreopsis maritima*, (melting point: 207-208°C; $[\alpha]_D^{25}:-98^\circ$), was synthesized by the condensation of 6,7-dihydroxy-chromene-(3)- β -D-glucoside with protocatechol aldehyde in acetic anhydride, followed by saponification. Orig. art. has: 2 formulas.

ASSOCIATION: Magyar Egyetem Szerves-Kozmai Tanoda, Budapest (Department of Organic Chemistry at the Technical University)

SUBMITTED: 17 Dec 64

ENCLs: 00

SUB CODE: GC, LS

MR REF Sov: 000

OTM: 009

JYB

Card 1/1 *TC*

L 01185-66

ACCESSION NR: AP5025820

HU/0005/65/071/006/0272/0273

AUTHOR: Farkas, Lorand; Pallos, Lasslo; Nogradi, Mihaly

18

8

TITLE: Aurones and aurone glucosides. Part 9: A new synthesis of leptosine

SOURCE: Magyar kemial folyoirat, v. 71, no. 6, 1965, 272-273

TOPIC TAGS: carbohydrate, plant chemistry, organic synthetic process

ABSTRACT: 6-hydroxy-7-methoxy-2-(3,4-dihydroxybenzal)-cumaranone-(3)- β -D-glucoside-(6), proven to be identical with leptosine, one of the glucosides in *Cota tinctoria*, was synthesized from 6-hydroxy-7-methoxycumaranone-(3)- β -D-glucoside-(6)-tetraacetate by condensation with protocatechol aldehyde in acetic anhydride, followed by saponification. The product had a melting point of 235-237°C.
Orig. art. has: 5 formulas.

ASSOCIATION: Műszaki Egyetem Szerves-Kemial Tanszéke, Budapest (Department of Organic Chemistry at the Technical University)

SUBMITTED: 17Dec64

ENCL: 00

SUB CODE: GC, LS

MR REF Sov: 000

OTHER: 009

JPRS

Card 1/1

L 33625-66

ACC NR: AP6025015

SOURCE CODE: HU/0005/65/071/011/0479/0481

AUTHOR: Farkas, Lorand; Pallos, Laszlo

19
CORG: Department of Organic Chemistry, Budapest Technological University, Budapest
(Budapesti Műszaki Egyetem, Szerves-Kémiai Tanszéke)TITLE: Aurones and aurone glucosides. X. Synthesis and final proof of structure of
bractein, a glucoside isolated from Helichrysum bracteatum (Vent.) Willd

SOURCE: Magyar környei folyoirat, v. 71, no. 11, 1965, 479-481

TOPIC TAGS: organic synthetic process, condensation reaction, plant chemistry

ABSTRACT: Condensation of 4,6-dihydroxycumaranone-(3)- β -D-glucoside(4)-tetra-acetate with gallaldehyde triacetate and subsequent saponification of the product yielded 4,6,3',4',5'-pentahydroxyaurone(4)- β -D-glucoside. This was found to be identical with the bractein isolated from Helichrysum bracteatum (Ven.) Willd. [JPRS: 33,906]

SUB CODE: 07. 06 / SUBM DATE: 17Mar65 / OTH REF: 007

LS

Card 1/1

084 1160

BARON, S., kand. fiz.-matem. nauk; PALLUM, E.; PETERSON, M.

On two theorems of Chow and their generalizations for double
series. Izv. AN Est. SSR. Ser. fiz. mat. i tekhn. nauk 11 no.4:
277-287 '62. (MIRA 16:1)

1. Tartuskiy gosudarstvennyy universitet i Institut kibernetiki
AN Estonskoy SSR.

(Series)

PALM, Adolf; ORA, A., red.; HINBERG, K., tekhn. red.

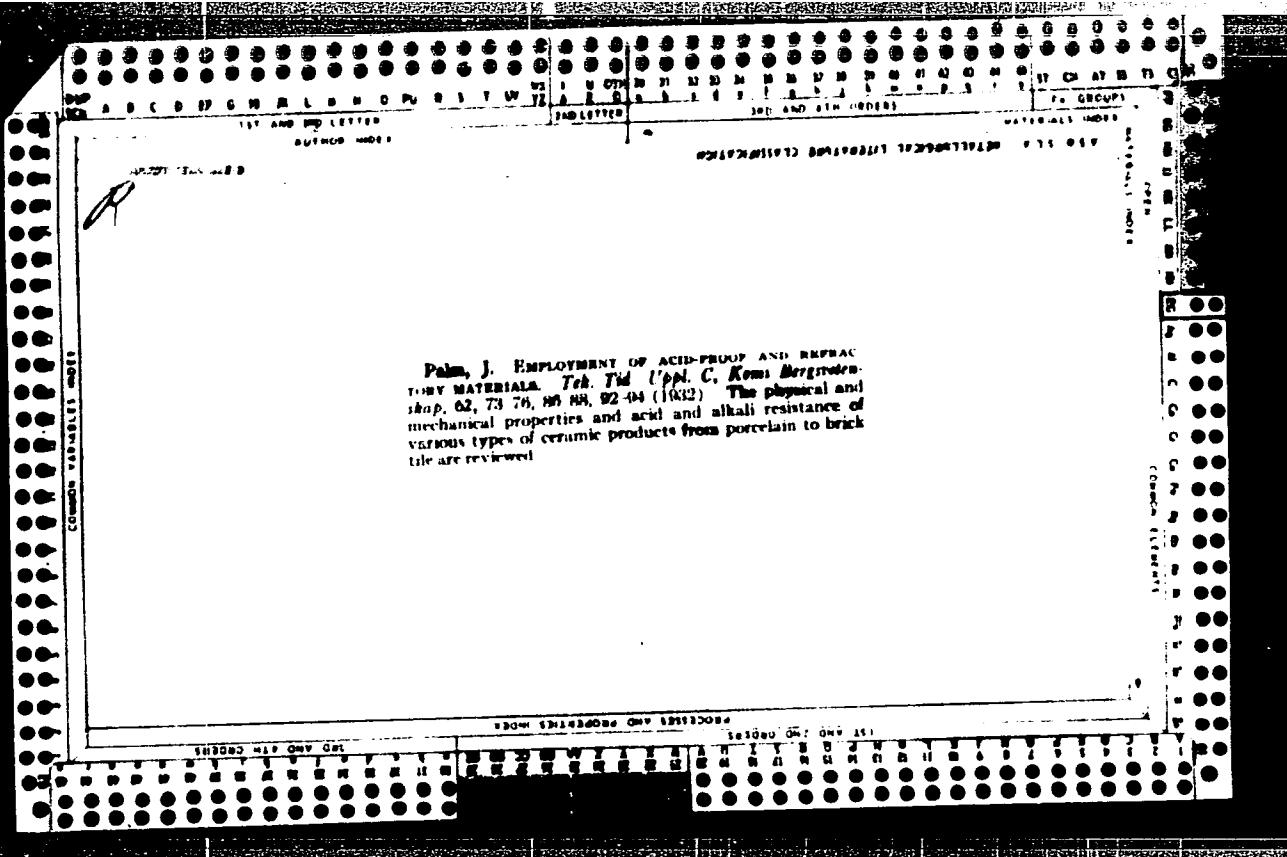
[Technical improvement of the SK-2,6 ensilage harvester]
Silokombaini CK - 2,6 tehnilisi täiustusi. Tallinn,
Eesti Riiklik Kirjastus, 1963. 33 p. (MIRA 17:1)

KASPIN, L.A., kand.ekonom.nauk; PAL'M, I.S., stershii nauchnyy sotrudnik;
KHORIKOV, A.N., stershii nauchnyy sotrudnik; SHEVCHUK, Yu.I.,
stershii nauchnyy sotrudnik; AKSENOK, D.G., inzh.; E'L'GOHT, Ye.G.
Prinimali uchastie: KARAKURCHI, M.I., kand.tekhn.nauk;
KUCHERENKO, K.R., kand.tekhn.nauk; PEDAN, M.P., nauch.sotr.; POPOV, V.Ye.,
nauchn.sotr.; GINZBURG, S.M., inzh.; SLIN'KO, B., red.; ZELENKOVA, Ye.,
tekhn.red.

[Economic aspects of the construction of four- and five-story
apartment buildings of large blocks of brick] Ekonomika vospovede-
niia 4-5 etazhnykh zhilykh zdanii iz kripayikh kirkichnykh blokov.
Kiev, Gos.izd-vo lit-ry po stroit. i arkhit. USSR, 1960. 112 p.
(MIRA 14:4)

1. Akademiya stroitel'stva i arkhitektury USSR. Institut organi-
zatsii i mekhanizatsii stroitel'nogo proizvodstva. 2. Sektor
ekonomiki stroitel'nogo proizvodstva Nauchno-issledovatel'skogo
instituta organizatsii i mekhanizatsii stroitel'nogo proizvodstva
Akademii stroitel'stva i arkhitektury USSR (for Kaspin, Pal'm,
Khorikov, Shevchuk, Aksenov, E'l'gort). 3. Nauchno-issledovatel'skiy
institut konstruktsiy (for Karakurchi, Kucherenko). 4. Glevkiyevstroy
(for Ginzburg). 5. Nauchno-issledovatel'skiy institut stroitel'nykh
materialov (for Pedan, Popov).

(Building, Brick)



PALM, L.

Introduce the use of the slide rule also in agriculture.

p. 371, (Sotsialistlik Põllumajandus) Vol. 12, no. 8, Aug. 1957, Tallinn, Estonia

SO: Monthly Index of East European Acessions (EAI) Vol. 5, No. 11 November 1957

LOOPMAA, V.; LOOG, P.; PAL'M, U. [Palm. U.]; PAST, V.; REEBEN, V.

Electronic potentiostat for electrochemical investigations.
Zhur. fiz. khim. 38 no.5:1374-1377 My '64. (MIRA 18:12)

I. Tartuskiy gosudarstvennyy universitet, laboratoriya
elektrokhimii. Submitted July 6, 1962.

Sov/70-13-1-1.

5(4)
AUTHORS:

Niklayeva, S. A., Pal'z, I. V.

TITLE:

Investigations of Electro-Capillary Curves in Mercury Containing Small Amounts of Copper (Isolated-vanadium-electrode method) (nykh krov'ykh na rtutti, sozvezdii-shey net plavlye v zolote i primesi met.)

PERIODICAL:

Zhurnal fizicheskoy khimii, 1959, Vol 33, No 1, p. 113

ABSTRACT:

Since metals used in practice always contain certain amounts of impurities due to other metals, an investigation of the effect of these impurities on the electro-capillary properties of the metal is of interest, and its electro-chemical behaviour is of interest. At present there are reports on this question by Rothmund (Rotmund) (Ref 1), Erdey-Grunz (Erdey) and Bajer (Bajor) (Ref 2), and by M. G. Smirnov, V. A. Smirnov and L. I. Antropov (Ref 3). The results of these investigations, however, do not agree well, it would seem, with each other. In the present case the investigations of the electro-capillary curves were carried out according to Gui-Lipmann's method (Refs 12,3). Mercury-amalgam was obtained electrolytically, and analyzed spectrographically after the experiment. It is the kafedra analiticheskoy khimii (Chair of Analytical Chemistry)

Card 1/2

SSV 74-73-11-1
Investigations of Electro-Capillary Curves in Mercury Containing Small Amounts of Copper

by the Head of the Laboratory Yu. P. Tsvetkov. The author presents curves of mercury and mercury-amalgam with 0.001, 0.002, 0.013% Cu in solutions of 0.1, 1, 2, 5, 4 and 1 N HCl, 0.1 and 1 N Hg_2SC_4 and 0.1 and 1 N Na_2SC_4 were investigated at room temperature. By the presence of copper and mercury the maximum of the electro-capillary curve is displaced in dependence on the electrolyte. A comparison of the experimental results with those of references 2 and 3 respectively shows that the differences occur because of the dependence of γ_{ex} on 2 factors: 1) the dependence on the potential of the point of the zero charge the value of which becomes more positive with the increasing purity of the mercury; 2) on the specific adsorption which is stronger in $HgCu$ -amalgam than in Hg . It increases in proportion with the copper content. There are 2 figures and 17 references, 6 of which are Soviet.

ASSOCIATION: Tartuskiy gosudarstvennyy universitet (Tartu State University),
Card 2/3

PAL'M, U.V.[Palm, U.]; PAST, V.E.[Past, V.]; REEBEN, V.A.[Reeben, V.]

Device for measuring the cathodic potential drop after cessation
of polarization. Zhur. fiz. khim. 35 no.5:1136-1139 My '61.
(MIRA 16:7)

1. Tartuskiy gosudarstvennyy universitet.
(Electromotive force)

PAL'M, U.V. [Palm, U.]; PAST, V.E.

Determination of the capacity of a smooth lead electrode by
measuring the potential drop. Dokl. AN SSSR 146 no.6:1374-1376
0 '62. (MIRA 15:10)

1. Tartuskiy gosudarstvennyy universitet. Predstavлено akademikom
A.N. Frumkinym.
(Electrodes, Lead) (Electromotive force)

PAL'M, G.V. [Palm, G.] ; IANT, V.E. [Bast, V.]

Hydrogen recombination with bismuth. Izv. Akad. Nauk SSSR, khim. 38 no. 3:
23.276 Mr. 1964
(MIRA 17:7)

I. Tsvetkov, G. Pal'm, V. Iant, Institute of Inorganic Chem., Kafedra neorganicheskoy khimii.

"APPROVED FOR RELEASE: 06/15/2000

CIA-RDP86-00513R001238910012-4

APPROVED FOR RELEASE: 06/15/2000

CIA-RDP86-00513R001238910012-4"

PAL'M, V., inzh.

Universal device for checking and straightening connecting rods.
Avt.transp. 38 no.6:54 Ja '60. (MIRA 14:4)

1. Tartuskiy avtoremontnyy zavod.
(Motor vehicles—Maintenance and repair)

"APPROVED FOR RELEASE: 06/15/2000

CIA-RDP86-00513R001238910012-4

PAL'M, V.A. inzh.

Attachment for boring thin-walled bushings. Avt.transp. 38 no.10:51
O '60. (MIRA 13:10)

(Lathes--Attachments)

APPROVED FOR RELEASE: 06/15/2000

CIA-RDP86-00513R001238910012-4"

PAL'M, V., inzh.

Attachment for grinding valve stems in circular grinding machines
Avt. tr. nsp. 4C no. 4:54 Ap '62. (GRA 1962)
(Grinding machines--Attachments)

PAL'M, V., Inc.

Sectional paint spray booth with filter. At 10:52 G (Z)
n. 10:52 G (Z)
(Painting Industrial-Equipment and supplies)

PALM, Viktor

Glycogen and fatt in the larval stage of Trematoda on the example
of Dolichosaccus rastellus and Haplometra cylindraceum (Plagiorchiida).
Acta parasit Pol 10 no.1/11:117-123 '62.

1. Parasitologische Abteilung des Bezirkss-Hygiene- Institutes
Potsdam, Kleinmachow b. Berlin Clara-Fetkin-Str. 23. Leiter:
Dr. H. Engelbrecht.

PALM, V.A.

Concerning some discrepancies in the theory of absolute reaction
rates. Zhur.fiz.khim 29 no.6:1116-1124 Je '55. (MLRA 9:1)

1.Tartuskiy gosudarstvennyy universitet.
(Chemical reaction, Rate of)

*P-1 m/f**Z**0
0**Chemical
PM's**AM SK*

✓ Problems in the activated-complex theory. V. A. Pol'm.
Voprosy Khim. Kinetiki, Kataliza i Reaktivnosti, Akad. Nauk S.S.R. 1955, 75-84. — Several inaccuracies are admitted in the "theory of abs. reaction velocities", derived from the theory of activated complexes that are used in most of the recent textbooks on phys. chemistry, chem. kinetics, and statistical mechanics, of which the principle is the assumption of an equil. concn. of activated complexes. It is shown how these inaccuracies can be relatively easily eliminated without interfering with the principles of the theory. No special assumptions are required in the activated-state concept in addn. or unknol. thermal, dynamic reactions. The existence of a "potential hole" at the energetic apex in exchange reactions may strongly affect the results, in comparison with values given by the theory. A coeff. of γ/α is introduced in the abs. velocity equation derived from the theory. W. M. Sorenson

PAL'M, V. A.

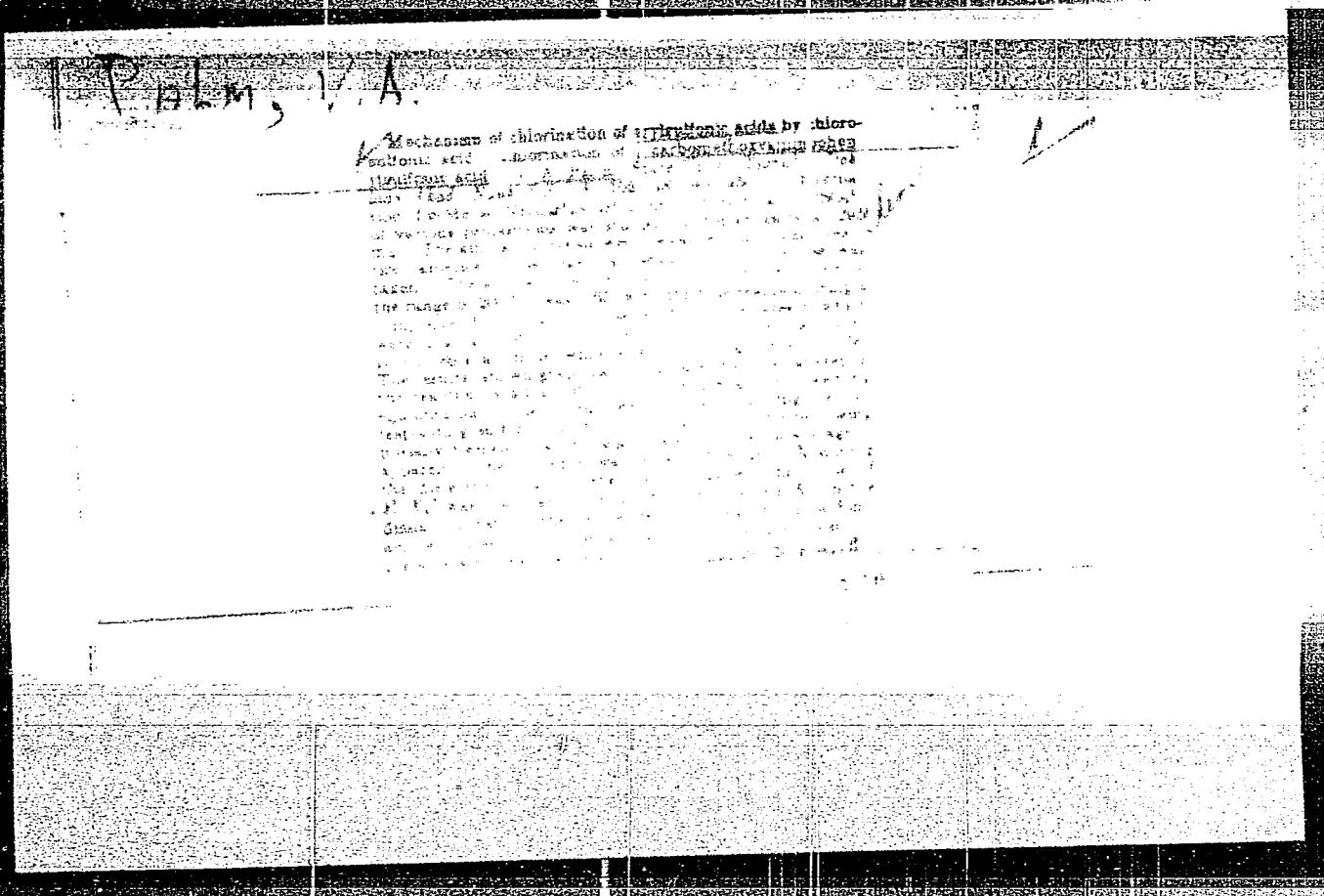
Acad Sci USSR. Inst of Chemical Physics

PAL'M, V. A.- "Investigation of the kinetics and mechanism of the reaction of chlorination of para-phenyl urethane sulfonic acid with chlorosulfonic acid." Acad Sci USSR. Inst of Chemical Physics. Moscow, 1955.
(Dissertation for the Degree of Candidate of Chemical Sciences.)

su: Kniahnaya Letopis' No. 13, 1956.

"APPROVED FOR RELEASE: 06/15/2000

CIA-RDP86-00513R001238910012-4



APPROVED FOR RELEASE: 06/15/2000

CIA-RDP86-00513R001238910012-4"

P. 1/1
USSR/Physical Chemistry - Solutions. Theory of Acids and Bases, B-11

Abst Journal: Referat Zhur - Khimiya, No 1, 1967, 45

Author: Pal'm, V. A.

Institution: Academy of Sciences USSR

Title: Acidity of the System $H_2SO_4-HSO_3Cl$

Original Periodical: Dokl. AN SSSR, 1969, Vol 108, No 2, 270-273

Abstract: The indicator-spectrophotometric method was used in measuring the acidity H_o of the system H_2SO_4 (I)- HSO_3Cl (II) over the composition range 0-100 mole percent II. The following indicators were used: n-nitrochlorobenzene ($pK = -11.49$), 2,4-dinitrotoluene ($pK = -12.6$) and n-nitrotoluene ($pK = -10.34$). As the basis of the calculation, the value of H_o at 100% H_2SO_4 is used; that value is -10.89. During the transition from 100% I to 100% II, H_o varies from -10.89 to -12.78, i.e., the acidity of II is 70 times greater than that of I.

Card 1/1

AUTHOR:

Pal'm, V. A.

76-32-5-1c/47

TITLE:

Investigation of the Chlorosulfonation of Aromatic Compounds
(Issledovaniye reaktsii khlorsul'firovaniya arokticheskikh
soyedineniy) III. The Influence of the Addition of Water and
 SO_3 on the Chlorination Velocity of the n-Phenylurethane-
 SO_3 sulfo Acid With Chlorosulfonic Acid (III. Vliyanije "otavok"
 SO_3 na skorost' khlorirovaniya n-feniluretilansul'fokis-
loty khlorsul'fonovoy kislotoj)

PERIODICAL:

Zhurnal fizicheskoy kemi, 1958, Vol. 32, Nr 5,
pp. 1030 - 1034 (USSR)

ABSTRACT:

In the further investigations of the dependence of the reaction
velocity constant on the composition of the reagent: the reaction
kinetics in the mixtures of H_2SO_4 and HSO_3Cl with small addi-
tions of water to mixtures of 0,5% oleum with chlorosulfonic
acid and oleum with different strength was investigated. The
experimental results obtained show that a heating or agitation
of the added amount of water effects a formation of HCl and
 H_2SO_4 from HSO_3Cl , with a complete reaction of the latter taking

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Investigation of the Chlorosulfonation of Aromatic Compounds. III. The Influence of the Addition of Water and SO_3Cl on the Chlorination Velocity of the n-Phenylurethanesulfo Acid With Sulfuric Acid

place at $120 - 140^\circ\text{C}$. It was observed that in this case the function of the acidity H_0 did not clearly determine the value of the reaction velocity, and that on the other hand the equation

$\lg K_1 = \text{const} - \alpha H_0$ expresses the function $\lg K_1$ of H_0 , with different values to be taken for the constant corresponding to the medium. Based on the observed phenomena the assumption of a parallel catalysis by protons and the SO_3Cl molecule was dropped as well as that based on the presence of two reaction mechanisms proceeding in parallel in anhydrous mixtures of H_2SO_4 - HSO_3Cl . There are 2 figures, 1 table and 5 references, 5 of which are Soviet.

ASSOCIATION: Gosudarstvennyy universitet Tartu (Tartu State University)
SUBMITTED: September 16, 1956
Carl Z/2 1. Phenylurethanesulfo acid--Chlorination
 2. Chlorosulfonic acid--Chemical reactions
 3. Water--Chemical reactions
 4. Sulfur oxide--Chemical reactions

AUTHOR: Pal'm, V. A.

76-32-²-19/43

TITLE: An Investigation of the Chlorosulfonation Reaction
(Issledovaniye reaktsii khlorsul'firovaniya).
II. The Kinetics of the Chlorination of p-Phenylurethane
Sulfo Acid by Chlorosulfonic Acid
(II. Kinetika khlorirovaniya n-feniluretilansul'fokisloty
khloresul'fonovoy kisloty)

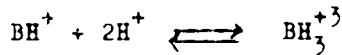
PERIODICAL: Zhurnal Fizicheskoy Khimii, 1958, Vol. 32, Nr 3,
pp. 620-627 (USSR)

ABSTRACT: Thanks to the determination of the absolute values of the acid function of the system $H_2SO_4-HSO_3Cl$ at various temperatures in an earlier paper, it is now possible to perform investigations of the function of the constant k_1 of the acid-function value at different temperatures. The method of preparation of the pure test substance is given, where a spectrophotometer of the type CF-4 was used in the observations of the course of reaction. With the use of the value for H_2O , under the conditions of the above-mentioned previous work, the function $\lg k_1$ of H_2O .

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An Investigation of the Chlorosulfonation Reaction. 76-32-3-19/43
II. The Kinetics of the Chlorination of p-Phenylurethane
Sulfo Acid by Chlorosulfonic Acid

at different temperatures can now be found. The author did not always obtain a straight line and the use of another value in the given case H^{+3} according to the scheme:



would be expedient, but it cannot be performed in default of a corresponding indicator. The kinetic curves obtained in the investigations approximately followed the molecular law when the changes of concentration of HSO_3Cl did not exceed 20% of the initial values. As a measure of the reactivity of any aromatic sulfo acid of the given type of reaction, a magnitude is used for whose calculation a mathematical derivation of the value for $\lg k_1$ yields a final formula according to which the obtained values are satisfactory, at medium concentrations for HSO_3Cl from 10 to 60 Mol%, whereas higher concentrations yield 2-3 times lower values for k_1 . The latter is explained by the fact

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An Investigation of the Chlorosulfonation Reaction. 76-32-3-19/43
II. The Kinetics of the Chlorination of p-Phenylurethane
Sulfo Acid by Chlorosulfonic Acid

that n-phenylurethane sulfo acid highly reduces the acidity and thus, also reduces the catalytic activity of the medium.

There are 3 figures, 3 tables, and 2 references, 2 of which are Soviet

ASSOCIATION: Tartuskiy gosudarstvenny universitet
(Tartu State University)

SUBMITTED: November 16, 1956

Card 3/3

AUTHOR: Pal'm, V. A.

76-32-2-22/38

TITLE: An Investigation of the Acidity of Chlorosulfonic Acid
(Issledovaniye kislotnosti khlorsul'fonovoy kisloty)
I. The Temperature Dependence of the Acidity of the
System $H_2SO_4-HSO_3Cl$
(I. Zavisimost' kislotnosti sistemy $H_2SO_4-HSO_3Cl$ ot temperatury)

PERIODICAL: Zhurnal Fizicheskoy Khimii, 1958, Vol. 32, Nr 2, pp. 380-387
(USSR).

ABSTRACT: The temperature dependence of the basicity constant of n-nitro-chlorobenzene was determined and the acidity function of the $H_2SO_4-HSO_3Cl$ -system was measured at various temperatures. The

magnitude $\lg \frac{[B]}{[BH^+]}$ was calculated from the measuring data

Card 1/5 of the absorption spectra of n-nitrochlorobenzene in the

An Investigation of the Acidity of Chlorosulfonic Acid
I. The Temperature Dependence of the Acidity of the
System $H_2SO_4-HSO_3Cl$

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corresponding mixtures of H_2SO_4 and HSO_3Cl , as described in reference 1. The linear dependence of:

$\lg \frac{[B]}{(BH^+)} \text{ on } \frac{1}{T}$ is satisfactory. In a 100% H_2SO_4 the acidity

function follows the equation $H_o = -\frac{6045}{4,574 T} - \frac{28,23}{4,574}$, when the

data of reference 3 are used. The values of $\lg \frac{[B]}{(BH^+)}$ at various

temperatures for n-nitrochlorobenzene in this medium are represented by the equation:

$\lg \frac{B}{BH^+} = -\frac{500}{4,574 T} + \frac{4,46}{4,574}$. Therefore the basicity of this

substance is marked by the following equation:

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An Investigation of the Acidity of Chlorosulfonic Acid
I. The Temperature Dependence of the Acidity of the
System $H_2SO_4^- - HSO_3Cl$

76-32-2-22/38

$$pK_a = - \frac{5540}{4,574 T} - \frac{32,70}{4,574}$$

The gradient of the straight of line

$$\lg \frac{[B]}{[BH^+]} \text{ denoting the dependence of } \frac{1}{T} \text{ on various } HSO_3Cl \text{ concen-}$$

trations is not the same. This points at the dependence of the transfer energy of the proton from the medium to the indicator on the composition of the $H_2SO_4^-$ and HSO_3Cl mixtures. The absolute values of the acidity function H_o of the $H_2SO_4^- - HSO_3Cl$ system at various temperatures were calculated according to the formula:

$$H_o = pK_a + \lg \frac{[B]}{[BH^+]}, \text{ where } pK_a \text{ is given by the equation (2)}$$

(see above). It is shown that the acidity of the $H_2SO_4^- - HSO_3Cl$

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An Investigation of the Acidity of Chlorosulfonic Acid
I. The Temperature Dependence of the Acidity of the
System $H_2SO_4-HSO_3Cl$

76-32-2-22/38

system decreases with the rise of temperature. In order to be able to interpret the data obtained the investigation is first of all directed towards the problem, on which magnitudes the values of the acidity function in any pure protonic acid in a mixture of a protonic acid and a base are dependent. It is shown that the values of the acidity function of a pure AH acid depend on the acidity constant as well as on the basicity constant of this acid. It is further shown that while the calculation of the acidity function of the $H_2SO_4-H_2O$ system without taking into account the ratio between the activity coefficients is possible in H_2SO_4 concentrations of more than 50 Mol% only, an analogous calculation of the $H_2SO_4-HSO_3Cl$ system can be carried out with sufficient exactitude for any ratio between these components. As regards H_2O and H_2SO_4 they form several chemical compounds with each other of which each must possess its own acidity- and basicity characteristics.

There are 3 figures, 4 tables, and 9 references, 4 of which are Soviet

Card 4/5

An Investigation of the Acidity of Chlorosulfonic Acid
I. The Temperature Dependence of the Acidity of the
System $H_2SO_4-HSO_3Cl$

76-32-2-22/38

ASSOCIATION: **Tartu State University**
(Tartuskiy gosudarstvennyy universitet)

SUBMITTED: November 16, 1956.

1. Nitrochlorobenzene--Properties 2. Nitrochlorobenzene--Temperature
factors 3. Nitrochlorobenzene--Spectra 4. Hydrogen ion
concentration--Determination

Card 5/5

5(4)

AUTHORS:

Tal'vik, A. I., Pal'm, V. A.

SOV/76-33-6-8/44

TITLE:

Investigation of the Kinetics and of the Mechanism of Acid Hydrolysis of Ethyl Acetate (Issledovaniye kinetiki i mekhanizma kislotnogo gidroliza etilatsetata)

PERIODICAL:

Zhurnal fizicheskoy khimii, 1959, Vol 33, Nr 6, pp 1214-1220 (USSR)

ABSTRACT:

Hammett (Ref 17) found a linear function between the logarithm of the reaction rate constant of the first order and the function of the acidity H_o of the medium in the acid hydrolysis of ethyl acetate (I), the inclination of this straight line being equal to one. On the strength of the foregoing, hydrolysis is preceded by a protonization of (I); the question, however, which is the oxygen atom to which the proton adds to effect hydrolysis, is still pending. Pal'm (Ref 18) derived a general equation setting a relationship between the reaction rate constant (RRK) of the acid-catalytic reaction and the acidity of the medium. Since in the (I) hydrolysis the addition of a single proton is sufficient, this equation is given the following simplified form:

Card 1/3

Investigation of the Kinetics and of the Mechanism of Acid SOV/T-35-6-8/44
Hydrolysis of Ethyl Acetate

$k = \frac{k_b}{K_a + h_o}$ (2) (K_a = constant of basicity of the reagent, $-\lg K_a^{H_2O}$).
With low h_o values this equation goes over to that of Hammett,
whereas in the case of higher h_o values the constant is equal to
the true RGK, which fact corresponds to a full protonization of
the reagent. The last mentioned concept is due to N. M. Chirkov,
who was the teacher of one of the authors of the present paper. In
the work under review, a more thorough investigation was carried
out of the mechanism of acid hydrolysis of (I), and the applicability
of equation (2) to this reaction was checked. A method was worked
out for the spectrophotometric control of the reaction course. A
table concerning the ultraviolet absorption spectra is supplied
(Table 1). Pertinent parallel experiments were made to control the
determination accuracy (Table 2); the data obtained for the
velocity constant k_1 are also given (Table 3). The invariability
of the value k_1 in the concentration range of 10 - 70% sulphuric
acid in which the quantities $[H_2C]$, $[H_3C^+]$, $[HSO_4^-]$ and $[H_2SO_4]$

Card 2/3

**Investigation of the Kinetics and of the Mechanism of Acid SNV/16-33-6-0/44
Hydrolysis of Ethyl Acetate**

change strongly is indicative of the fact that the reaction stage determining the rate of hydrolysis, is really monomolecular. A previous reaction mechanism is suggested for the (1') hydrolysis, and the applicability of reaction (2) to this reaction is confirmed. There are 3 figures, 3 tables, and 21 references, 4 of which are Soviet.

ASSOCIATION: Tatsinskij's Jarunverksgesellschaft (East German Company)

SUBMITTED: September 5, 1957

Card 3/3

5(4)

AUTHORS: Khaldna, Yu. L., Tal'vik, A. I.,
Pal'm, V. A.

SOV/20-126-1-32/62

TITLE: The Dependence of the Rate of Acidic-catalytical Reaction on
the Basicity of the Reagent in the Case of the "General
Acidic Catalysis" (Zavisimost' skorosti kislotno-kataliches-
koy reaktsii ot osnovnosti reagenta v sluchaye "obshchego kis-
lotnogo kataliza")

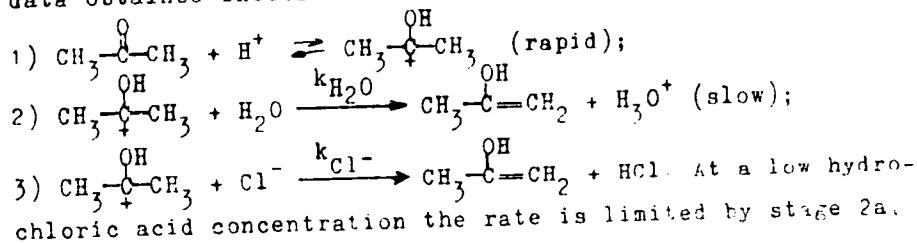
PERIODICAL: Doklady Akademii nauk SSSR, 1959, Vol 126, Nr 1, pp 119-127
(USSR)

ABSTRACT: The conception of "general acidic catalysis" is defined as
a reaction, the rate constant of which is representable by
means of a polynomial, in which every term is proportional
to the concentration of any acid existing in the system.
In concentrated acidic solutions and in some water-free
acids the logarithm of the rate constant in the case of many
acidic-catalytical reactions depends linearly on the acidity
function H_0 of L. P. Hammet (Refs 2,3). The development of
the reaction may be represented by the scheme (3)
1) $B + H^+ \rightleftharpoons BH^+$ (rapid); 2) $BH^+ + C \xrightarrow{k_0}$ reaction prod-
uct (slow). In order to find out whether this scheme applies

Card 1/3

The Dependence of the Rate of Acidic-catalytical Reaction on the Basicity of the Reagent in the Case of the "General acidic Catalysis" SCV 30-126-1-32-62

to all general acidic catalyses, the kinetics of the enolization of acetone in an aqueous hydrochloric acid solution with concentrations of from 0.04-11.2-normal was investigated at 15, 25, and 35°. The reaction rate was measured spectrometrically according to the decrease of the bromine content by brominating the acetone. The reaction rate of bromation is, according to reference 6, equal to that of enolization (Fig. 1). In the case of very high concentrations of hydrochloric acid, the rate constant is proportional to the concentration of the chloric ion and depends no longer on acidity (Table 1). The data obtained indicate the following mechanism:



Card 2/3

The Dependence of the Rate of Acidic-catalytical Reaction on the Basicity of the Reagent in the Case of the "General Acidic Catalysis" SCV 2 -12t-1-32 62

and at high concentration by stage 2b. It is concluded here-
from that in the case mentioned the so-called "general
acidic catalysis" does not differ basically from "specific
acidic catalysis", in which the rate constant is proportional
to the concentration of the hydrogen ion. The rate of the
reaction is limited by the concentration of the protonized
form of the reagents. The catalytic activity of the medium
is proportional to its acidity and not to the concentration
of an arbitrary acid. Thus, the reaction mechanisms sug-
gested in references 7-11 are refuted. There are 1 figure,
1 table, and 19 references, 5 of which are Soviet.

ASSOCIATION: Tartuskiy gosudarstvennyy universitet (Tartu State University)
PRESENTED: December 24, 1958, by V. N. Kondrat'yev, Academician.
SUBMITTED: December 11, 1958

Card 3/3

5(4)

AUTHORS: Palm, V. A., Kõrgesaar, A. O. SOV/20-127-2-40/70

TITLE: On the Problem of the Mechanism of Reaction Between Alcohols and Halogen Hydracids

PERIODICAL: Doklady Akademii nauk SSSR, 1959, Vol 127, Nr 2, pp 380-383 (USSR)

ABSTRACT: To clarify the reaction mentioned in the title, an investigation was made of the chlorination of absolute ethanol as well as of two different aqueous ethanol solutions in the temperature range 100 - 140°C. Table 1 gives the values of the molecular constant k_1^o of the reaction rate for HCl concentrations $> 0.1 \text{ mol/l}$, table 2 the values for $\text{HCl} < 0.1 \text{ mol/l}$. Reaction is monomolecular at $\text{HCl} > 0.1 \text{ mol/l}$. With $\text{HCl} < 0.1 \text{ mol/l}$ the order of reaction rises. Table 3 shows the electrical conductivity of the HCl solutions in absolute ethanol at 101° in a concentration range of from $3.2 \cdot 10^{-4}$ to 0.5 mol/l. Proceeding from C. N. Hinshelwood's assumption (Ref 4) that in slow reaction stages dehydration of ion pairs occurs, the reaction constant k_1 for

Card 1/3

On the Problem of the Mechanism of Reaction Between
Alcohols and Halogen Hydracids

SOV/20-127-2-40/70

low HCl concentrations is computable according to the following formula based on the dependence of the concentration of the ion pairs on the HCl concentration: $k_1 = k_1^0 \frac{1 - \lambda/\lambda_\infty}{1 - \lambda_0/\lambda_\infty}$ (5).

Herein, λ/λ_∞ is the mean value of the dissociation degree of the ion pairs in the concentration range of HCl, in which k_1^0 was measured. As is shown in table 2, the values obtained with (5) are in good agreement with experimental data. This confirms Hinselwood's hypothesis. Table 1 reveals that water added to ethanol inhibits chlorination. This is explained by the inhibiting effect of basic additions to the reaction of acid catalysis. The protonized particles (ethoxonium ions) enter reaction with water, involving a drop in their concentration and consequently also of the reaction rate. The following equilibrium is brought about: $H_3O^+ + C_2H_5OH \rightleftharpoons C_2H_5OH_2^+ + H_2O$.

The concentration constant of this equilibrium was measured by means of p-nitroaniline as indicator in the temperature range of from 15 to 45° in 96.6% ethanol. The values of K obey to the

Card 2/3

On the Problem of the Mechanism of Reaction Between
Alcohols and Halogen Hydracids SOV/20-127-2-40/70

equation $\lg K = 7.87/4.574 - 5100/4.574T$ (7). The values extrapolated therefrom for the temperature range 110 - 130° (Table 4) are in good agreement with experimental data. If dehydration of the ion pairs is the stage limiting the reaction rate, the reaction between HBr and ethanol is expected to occur more quickly in consequence of the high nucleophilic nature of Br⁻. This is confirmed by provisional data. The authors thank the student L. Ploom for assistance in the experimental part. There are 4 tables and 9 references, 3 of which are Soviet.

ASSOCIATION: Tartuskiy gosudarstvennyy universitet (Tartu State University)
PRESENTED: March 5, 1959, by V. N. Kondrat'yev, Academician
SUBMITTED: March 3, 1959

Card 3/3

KHALDINA, Yu.L. [Haldna, J.]; PAVL'IN, V. N.

Problem on the reduction of basicities to a single standard state.
Study of the basicity of acetone. Dokl. AN SSSR 135 no.3:667-670
N '60. (MIFR 13:12)

1. Tartuskiy gosudarstvennyy universitet. Predstavлено akad. V.N.
Kondrat'yev. (Act ne) (bases (Chemistry))

PAL'M, V.A.

Structure and reactivity of organic compounds (quantitative correlations).
Usp.khim. 30 no.9:1069-1123 S '61. (MIA 14:10)

1. Tartuskiy gosudarstvennyy universitet.
(Chemistry, Physical and Theoretical)
(Chemical structure) (Organic compounds)

KOZLOVA, Kh.R.; PAL'M, V.A.

Synthesis of tetraphenyl boron sodium in a tetrahydrofuran medium.
Zhur. ob. khim. 31 no.9:2922-2923 S '61. (MIRA 14:9)

1. Tartuskiy gosudarstvennyy universitet.
(Boron compounds) (Furan)

PAL'M, V.A. [Palm, V.]; KHALDNA, Yu.L. [Kaldna, J.]; TAL'VIK, A.I.
[Talvik, A.]; MEY, A.F. [Mei, A.]

Protonation of carbonyl compounds and the mechanism of the
acid hydrolysis of esters. Zhur. fiz. khim. 36 no.11.
2499-2501 N'62. (MIFI A 12:5

1. Tartusskiy gosudarstvennyy universitet.

PAL'M, V.A.; VIZGERT, R.V.

Correlation of the rate constants in the alkaline hydrolysis
of aromatic sulfonic acid esters as dependent on their structure.
Dokl. AN SSSR 142 no.5:1091-1094 F '62. (MIRA 15:2)

1. Tartuskiy gosudarstvennyy universitet L'vovskiy politekhnicheskiy
institut. Predstavлено академиком V.N.Kondrat'yevym.
(Sulfonic acid)
(Hydrolysis)

PALM, V. p. red.

"Transactions of the Conference on the problems of the use of Correlation Equations in Organic Chemistry. Trudy po problemam primenija korrelatsionnykh uravnenij v organicheskoi khimii. Tartu. Tartuskiy gos. univ. Vol. 1. 1962. 141 p." MIR 1964.

1. Konferentsiya po problemam primenija korrelatsionnykh uravnenij v organicheskoy khimii. Tartu. 1962.

ESILEV, Anatoly Vasil'evich, doctor of phys., 1915; R. G. M. S., member of the Komsomol, 1943; M. V. Lomonosov Moscow State University, professor; 1948; V.A. Ulyanov, participant; Andrei Ilyin, phys., teacher, music, 1948; Shchukin, Irina, 1948.

[Redacted] [Redacted] [Redacted] [Redacted]
[Redacted] [Redacted] [Redacted] [Redacted]
[Redacted] [Redacted] [Redacted] [Redacted]

Leningradskiy gosudarstvennyy universitet filologii i AN SSSR (for party).

AERAMOVA, N.A., nauchn. sotr.; SEL'CHENKO, G.V., kand. tekhn. nauk; BEREBLIT, V.V., nauchn. sotr.; VASIL'YEV, V.P., kand. khim. nauk; DOLYGIN, D.P., doktor khim. nauk; IOFFE, S.V., dokt. khim. nauk; KAYRASKIY, Yu.L., nauchn. sotr.; KARPOVA, I.P., kand. khim. nauk; KOFYLEN, S.A., doktor khim. nauk; LUTUGINA, N.V., kand. khim. nauk; MATERCOVA, Ye.4., kand. khim. nauk; MORACHEVSKIY, Al.G., kand. khim. nauk; NIKEEV, A.E., kand. khim. nauk; PAL'Y, V.A., kand. khim. nauk; RAINOVISH, V.A., kand. khim. nauk; SOKOLOV, I.I., kand. khim. nauk; FRIDRIKHEBERG, E.A., kand. khim. nauk; TSYGLI, Ye.N., nauchn. sotr.; SHAGITSULTANOVA, G.A., kand. khim. nauk; SHKODIN, A.M., doktor khim. nauk; YATSIMIRSKIY, K.S.; GRIGOROV, O.N., doktor khim. nauk, red.; ZASLAVSKIY, A.I., kand. khim. nauk, red.; MORACHEVSKIY, Yu.V., prof., red.; RABINSKIY, F.Yu., kand. khim. nauk, red.; POZIN, E.Ye., doktor tekhn. nauk, red.; POLOAY-KOZHILS, S.A., doktor khim. nauk, red.; PRISTASOV, A.N., kand. fiz.-mat. nauk, red.; ROMANKOV, F.S., red.

(Handbook for the chemist) Spravochnik khimika, 2. izd., perer. i dop. Moscow, Khimiya. Vol.3. Publ. 1964 p. (GIA 18:1)

i. Chlen-korrespondent AN SSSR (for Romankov). . . Deputativnyy chlen Ak. Ukr. SSR (for Yatsimirskiy).

"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001238910012-4

APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001238910012-4"

PAL'M, YA.M.

"Built-up Geodetic Signals From Concrete Construction," by
P. A. Sokolov and Ya. M. Pal'm, Tr. Novosibir. in-ta. inzh.
geod., aerofotos'yemki i kartogr., No 7, 1956, pp 21-28
(from Referativnyy Zhurnal -- Astronomiya, Geodeziya, No 2,
Feb 57, Abstract No 1642)

To save structural wood and rolled steel in industrial and forestless regions, the use of built-up geodetic signals with instrumental pyramids of concrete girders and with demountable wooden or metallic external pyramids is suggested. Examples are given for computing 3- or 4-faced pyramids, and for construction of a Duran to the 12-m signal, and drawings are included.
(U)

sum. 1360

PAL'M, YU. A.

"Some Problems of Stability of Longitudinally Compressed Bars. Elastic and Plastic Ranges)." Min Higher Education USSR, Saratov Automotive and Road Inst imeni V. M. Molotov, Saratov, 1955. (Dissertation for the degree of Candidate of Technical Sciences)

SO: N-972, 20 Feb 56

PAL'M, Yu.A., inzh.; TARASOV, V.A., inzh.; SHMAKOV, Yu.M., inzh.

Experimental investigation of stiffener deformations on an
inclined shell of double curvature. Mat. po met. konstr.
no.7:163-176 '62. (MIRA 1":1)

VAKHURKIN, V.M.; GLADSHTEYN, L.I.; KARMILOV, S.S.; KLIMOV, S.A.;
LEVITANSKIY, I.V.; MALININ, B.N.; NOSOV, A.K.; PAL'M,
Yu.A.; POLYAK, V.S.; PCPOV, G.D.; RASSUDOV, V.M.;
KRASYUKOV, V.P.; SOKOLOV, A.G.; Prinimali uchastiye:
GORBATSKIY, Ye.I.; MATVEYEV, S.S.; STRELETSKIY, N.S.,
prof., retsenzent; MUKHANOV, K.K., dots., retsenzent;
BOLOTINA, A.V., red.; MIKHEYEVA, A.A., tekhn. red.

[Light-weight supporting metal structures] Oblegchennye
nesushchie metallicheskie konstruktsii. Moskva, Gos-
stroizdat, 1963. 282 p. (MIRA 17:2)

PAL'MA, I.S., starshiy nauchnyy sotrudnik

Using correlation methods in investigating economic activities
of building organizations. Trudy MIEI no.14:559-567 '59.
(MIRA 13:1)

1. Nauchno-issledovatel'skiy institut organizatsii, mekhanizatsii
i ekonomiki stroitel'stva Akademii stroitel'stva i arkhitektury
USSR.

(Construction industry--Costs)

PALIMA IS.

25(5)150(5) **PLATE I: WORK INFORMATION** 257/253
Source: Industrial and Economic Survey, Local Survey Organization.
 Foreign participation in industrial enterprises, especially in capitalistic type production, is one of the main problems of the present-day economic situation. The problem of increasing foreign participation in the economy has been approached by the Government of India in its Second Five-Year Plan (1956-61). According to the Second Five-Year Plan, 20,000 enterprises will be established by foreign investors. The following table gives the estimated number of enterprises to be established by foreign investors in different industries.

This collection of articles is intended for staff members of construction organizations, design bureaus, and scientific research establishments as well as for family members and students of institutions of higher education.

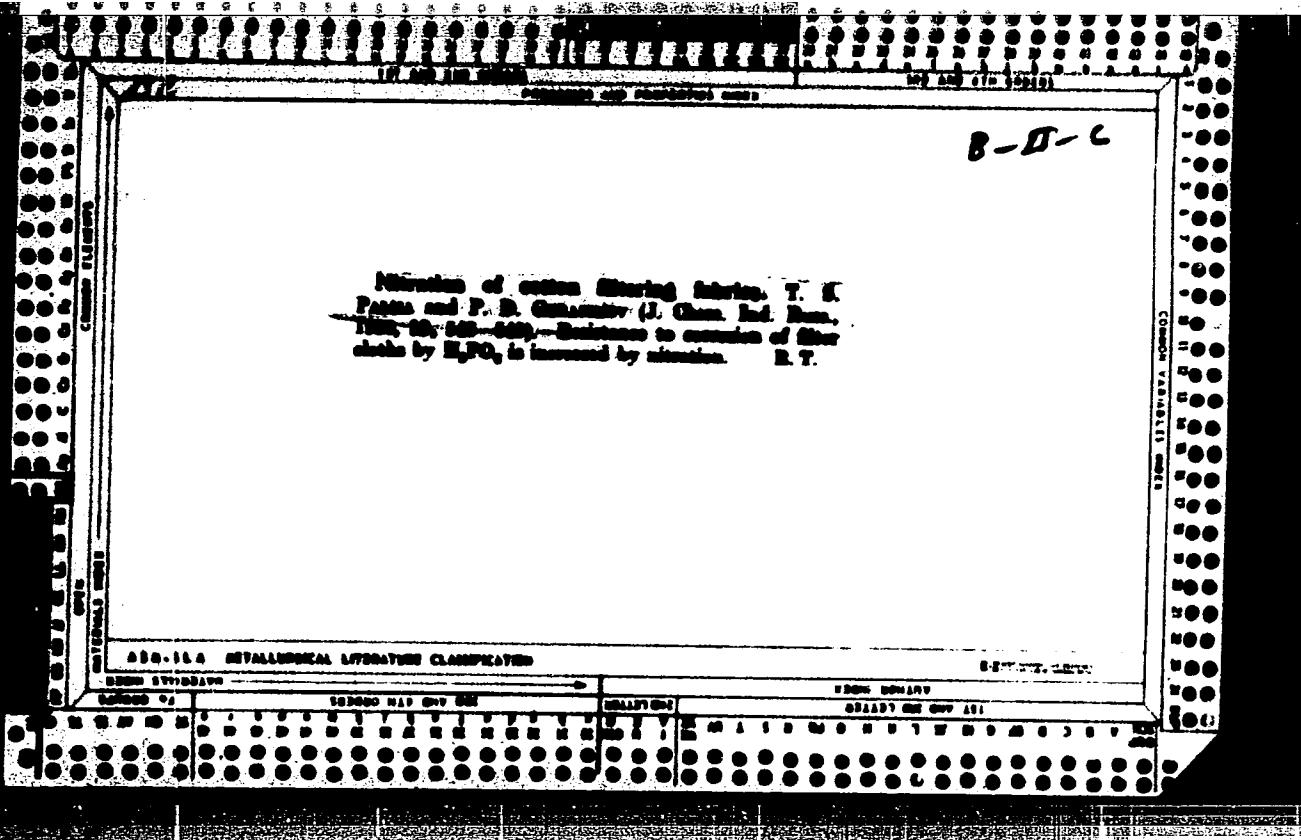
This collection of reports on construction problems was originally presented and discussed at a scientific-technical conference held in Moscow in February 1958 under the auspices of the All-Union Society of Engineers and Architects, the All-Union Society of Economic and Economic Benefits from Scientific Organizations, Possibilities of organizing and planning construction projects by improving methods of effort, by construction and design organizations to reduce the costs of construction and building operations, to distribute economic accountability among all participants in construction, to increase the productivity of labor, and to boost work and planning efficiency. The problems of preparing estimates, making financial forecasts, and financing construction projects are discussed. No references are given.

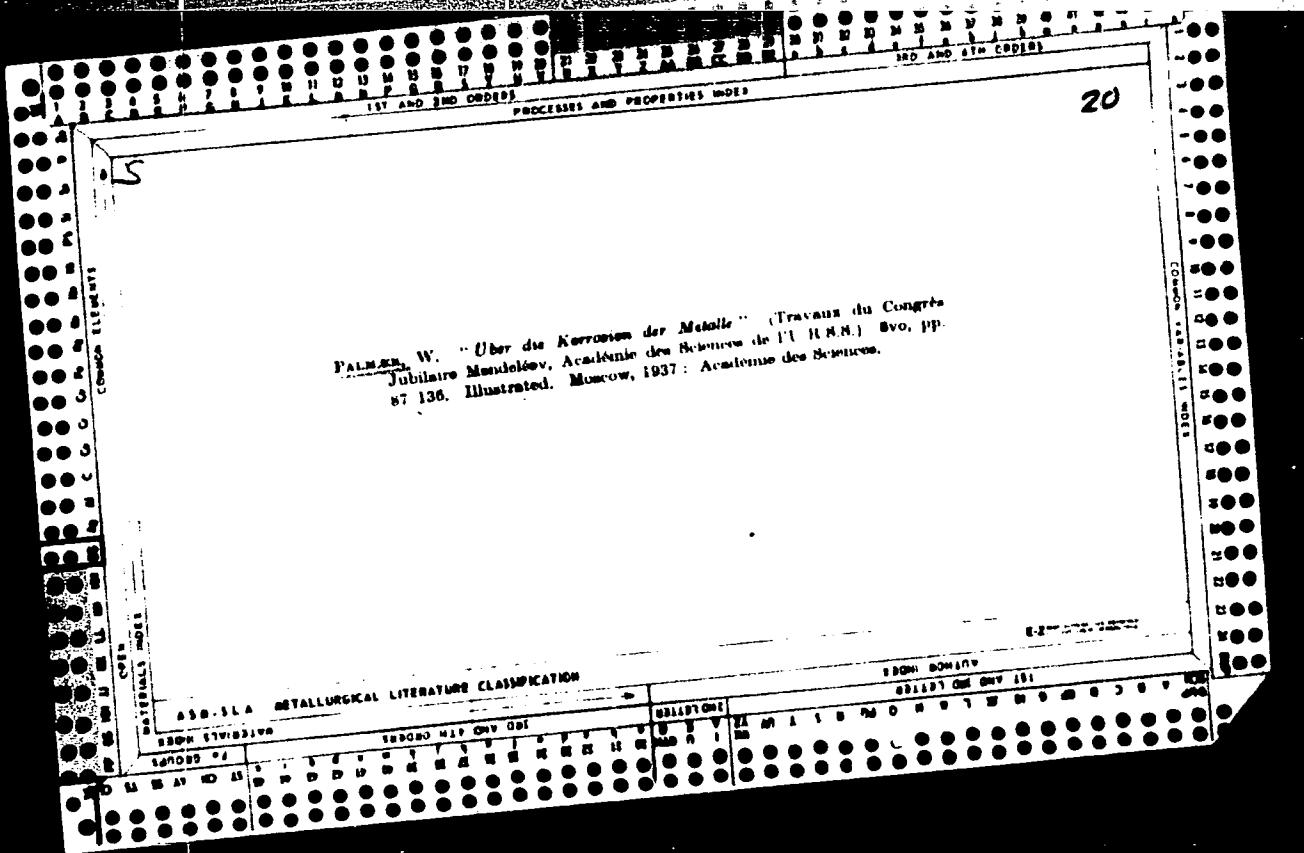
<u>Barney, T. S., and M. I. Mandelkern.</u> Peculiar Features and Scope of Engineering Construction Cost Estimates	509
<u>Barker, J. A.</u> Development of Technical Standardization and Possibilities of Improving Technical Standards	511
<u>Bashaw, T. A.</u> Price Calculation Problems	517
<u>Bellman, B. E.</u> Improvement of Standardized Estimates and the Pattern of Pricing Estimates for Construction Products	522
<u>Bendheim, D. H.</u> Possibility of Lowering Costs of Prefabricated Reinforced Concrete Structures	529
<u>Bellamy, L. F.</u> Application of Correlation Methods to the Study of the Economic Activities of Construction Organizations	539
<u>Long, B. T.</u> Economic Accountability in Construction Works	563

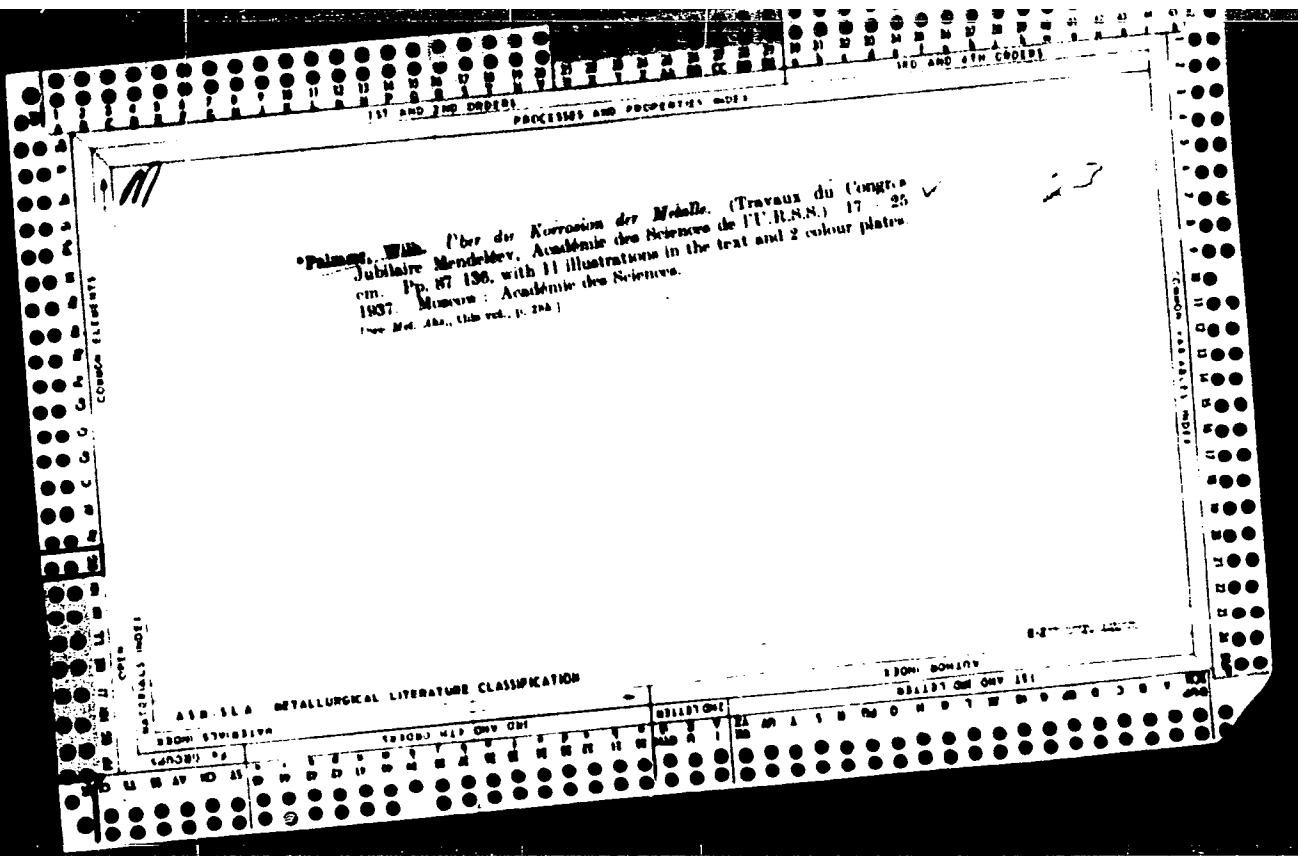
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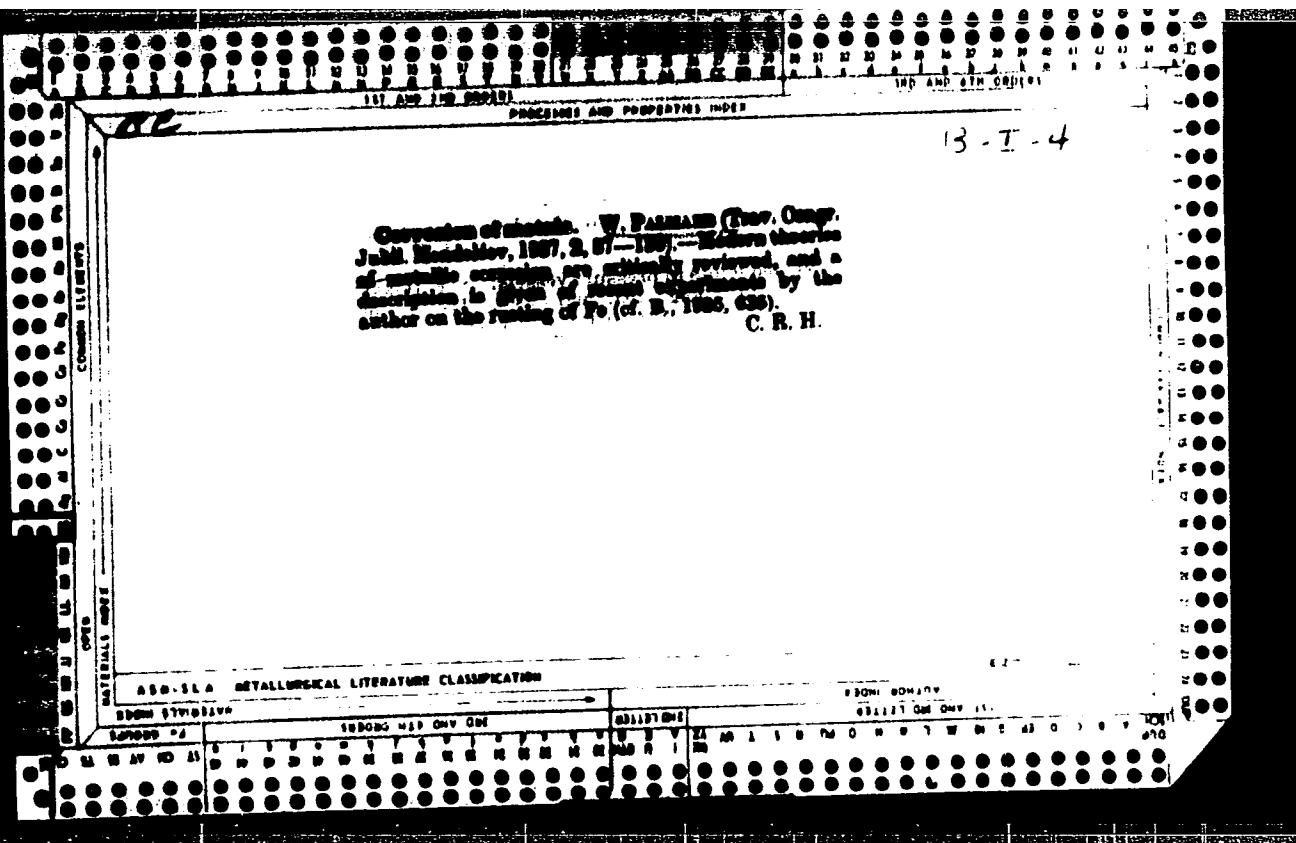
APPROVED FOR RELEASE: 06/15/2000

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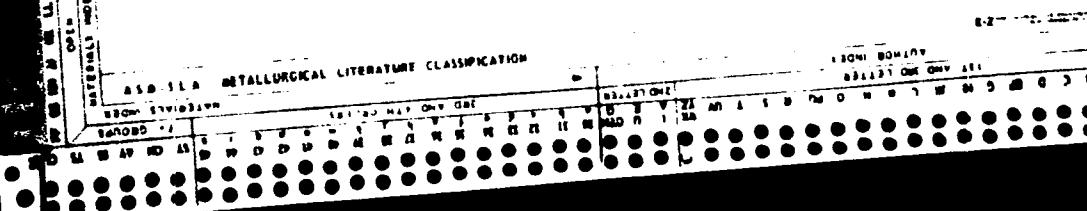




Ca

Working up bast fibers. Antal, Buday-Goldberger
and Gustáv Palme. Hung. 132,345. Dec. 15, 1943.
Scutched-green flax, flax wastes, Mellotex, A.
cotton, or lin straw) are treated in the same container
two or more times with a boiling soln. contg. 10-20 g/l
NaOH for 4 hrs. at 110-115° and 2-8 atm. pressure, then
further boiled with oxidizing agents (e.g., 1-2% of a 30%
soln. of H₂O₂) under the same conditions. After me-
chanical treatment the fibers are worked up as usual.
István Fájdy

25



PALMAI, M.

PALMAI, M. Outline of the geography of the soil's surface in the Szeged area.
p. 77.

Vol. 4, no. 1, 1955
FOLDRA JZI ERTESTITO
GEOGRAPHY & GEOLOGY
Budapest, Hungary

So: East European Accessions, Vol. 5, no. 5, May 1956

PALMAI, M.

Drinking-water supply of the city of Szeged. p. 71.

FOLDRAJZI ERTESENTO. (Magyar Tudomanyos Akademia. Foldrajztudomanyi Kutatcsoport)
Budapest, Hungary. Vol. 8, no. 1, 1959.

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unclu.

LEKA, Tivadar; PALMAI, Zoltan

Performance tests on cutting machines at the Ozd Metallurgical works.
Munka szemle a műszaki munkában.

IPAK, Tivadar, i ALMAI, Zoltan

Capacity test of roll tapes by recording instruments. Roll
szemle 7 no. 714-19 31'6".

PALMAJ, Stevan

Mechanized mot ~~ba~~ - ~~asino~~gradnja 5 no.2:34-35
Jl '62.

PALMAN, D.

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1. - P/W

Palman, Dominik. Über eine räumliche kubische Inversion und einige ihrer Ergebnisse. Bull. Internat. Acad. Yugoslav. Cl. Sci. Math. Phys. Tech. (N.S.) 12, 103-107 (1954).

Palman, Dominik. On a spatial cubic inversion and certain of its uses. Rad Jugoslav. Akad. Znan. Umjet. Odjel Mat. Fiz. Tehn. Nauke 296, 199-214 (1953). (Serbo-Croatian)

Serbo-Croatian version of the preceding paper.

Geometry

Handwritten notes:

PAINTER, J.

Surfaces of the 30 carrier with supply dump site, 1000' above sea level, 1000' below sea level
CC: Monthly List of East European Accessions, FBI, DC, Tel. 202-540-5000, 1970

PALMAN, D.

Entire circular curves of the third order in a hyperbolic plane. Bul sc Youg 8 no. 1/2: 19 F-Ap '63.

1. "Elektroprojekt", Zagreb.

PALMAN, Dominik, (Zagreb)

Constructing onefold focuses of the 3d-order-circular curve. In
German. Gl mat fiz Hrv 15 no.3:189-197 '60. (KEAI 10:8)
(Curves) (Circle)

PALMAN, Dominik, (Zagreb)

Projective metric and isometric transformations on the 2d order
planes. In German. Gl mat fiz Hrv 15 no.3:199-220 '60.

(EEAI 10:8)

(Isometric projection) (Transformations(Mathematics))
(Curves)

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DATE 06-15-2000 BY SPK

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HEREIN IS UNCLASSIFIED DATE 06-15-2000 BY SPK

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CIA-RDP86-00513R001238910012-4"

FRIYMAK, A.K., doktor sel'khoz. nauk, red.; KOLESNIKOV, M.A., kand. sel'khoz. nauk, red.; TRUSEVICH, G.V., kand. sel'khoz. nauk, red.; PAL'MAN, V.I., red.; ZANADVOVKOV, S.M., red.; RUCH'YEV, L.I., tekhn. red.; KHLOBORDOV, V.I., tekhn. red.

[Research achievements of State Northern Caucasus Regional Research Institute of Fruit Culture and Viticulture] Itogi nauchno-issledovatel'skoi raboty. Krasnodar, Krasnodarskoe knizhnoe izd-vo, 1959. 335 p.
(MIRA 14:7)

1. Severo-Kavkazskiy zonal'nyy nauchno-issledovatel'skiy institut sadovodstva i vinogradarstva.

(Caucasus, Northern--Fruit culture) (Caucasus, Northern--Viticulture)

PAL'MAN, Ye., inzh.

Machine tool for shaped cutting of pipe. Prom.stroi.i inzh.
soor. 4 no.2:59-61 Mr-Ap '62. (MIRA 15:11)
(Pipe cutting)

KIĆIĆ, M., profesor, dr.; PAPO, profesor, dr.; ANTIC, M., docent, dr.;
BERVAN, M., dr.; KIĆIĆ, R., dr.; JANJIC, M., dr.; PALMAR, I., dr.

Current problems in the treatment of hyperthyroidism. Vojno-sanit. pregl. 21 no.10:600-613 0 '64

IVANOV, Aleksandar, sanitetski potpukovnik dr; MORELJ, Marja, general-major sanitetske slusbe doc. dr; PALMAR, Ivan, sanitetski potpukovnik dr; VUKS, Ljubomir, sanitetski pukovnik prof. dr.

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